**Answers to problems in page 327**

## Answers to Practice Problem J

1. 4.01 g $CO_2$

## Homework

GENERAL

**Additional Practice** Racecars often burn ethanol, $C_2H_5OH$, for added performance. The products are carbon dioxide and water. The density of ethanol is 0.816 g/mL, the density of $CO_2$ is 1.997 g/L, and the density of $O_2$ is 1.331 g/L.

1. If the car holds $1.00 \times 10^5$ mL of ethanol and all of the carbon in it forms carbon dioxide, what volume of carbon dioxide is added to the air? Ans. $7.81 \times 10^4$ L $CO_2$
2. What volume of oxygen is needed to react with 78.3 L of ethanol? Ans. $1.00 \times 10^5$ L $O_2$

LS Logical