Answers to problems in page 307.

Answers to Practice Problems B

1. 45.6 g Al
2. 44.6 g Al_2O_3
3. 679 g Fe_2O_3
4. 107 g Fe

Homework — GENERAL

Additional Practice

1. What mass of H_2O is produced if 65.2 g $CaCO_3$ reacts with excess H_3PO_4, to form $Ca_3(PO_4)_2$, H_2O, and CO_2? Ans. $3CaCO_3(s) + 2H_3PO_4(aq) \rightarrow Ca_3(PO_4)_2(s) + 3H_2O(l) + 3CO_2(g)$; 11.7 g H_2O
2. What mass of O_2 forms when 49.89 g $KClO_3$ decomposes? (KCl also forms.) Ans. $2KClO_3(s) \rightarrow 2KCl(s) + 3O_2(g)$; 19.54 g O_2
3. What mass of ammonia is formed when 7.50 g N_2 reacts with excess H_2? Ans. $N_2(g) + 3H_2(g) \rightarrow 2NH_3(g)$; 9.12 g NH_3

LS Logical