Answers to problems in page 252-253.

42. a. $127 \text{ g } I_2$
    b. 675 g PbS
    c. $233 \text{ g } C_4H_{10}$
    d. $103 \text{ g } Al_2(SO_4)_3$
    e. $35.4 \text{ g } CuSO_4$
43. a. 58.44 g NaCl
    b. $36.04 \text{ g } H_2O$
    c. $260 \text{ g } Ca(OH)_2$
    d. $163 \text{ g } Ba(NO_3)_2$
44. 3.24×10^{22} atoms Au
45. 1.337×10^{24} formula units $ZnCl_2$
46. 9.396×10^{21} molecules naphthalene
47. 2.79×10^{24} atoms Al
48. $5.53 \text{ mol } H_2O$
49. a. $4.99 \times 10^{-2} \text{ mol } (NH_4)_2SO_4$
    b. $61 \text{ mol } Ca(OH)_2$
    c. $7.49 \times 10^{-2} \text{ mol } H_2SO_4$