

Answers to problems in page 180

Answers to Practice Problems A

- a. $\text{Ca}(\text{CN})_2$
- b. $\text{Rb}_2\text{S}_2\text{O}_3$
- c. $\text{Ca}(\text{CH}_3\text{COO})_2$
- d. $(\text{NH}_4)_2\text{SO}_4$

5. a. calcium nitrite
b. iron(III) hydroxide

- c. ammonium dichromate
d. copper(I) acetate

6. a. Na_2O c. Ag_2S
b. Mg_3P_2 d. NbCl_5

7. a. rubidium oxide
b. iron(II) fluoride
c. potassium nitride

8. a. HgSO_4 c. $(\text{NH}_4)_3\text{PO}_4$
b. $\text{Li}_2\text{S}_2\text{O}_3$ d. KMnO_4

Quiz GENERAL

1. What does a Roman numeral after the name of a cation indicate? **Ans.** The charge on the cation.
2. Why must the number of positive charges equal the number of negative charges in all ionic compounds? **Ans.** Because all compounds must be neutral in charge.
3. What is the name of SrCrO_4 ? **Ans.** Strontium chromate
4. How many oxygen atoms are there in $\text{Mg}(\text{NO}_3)_2$? **Ans.** six
5. What is the formula for rubidium phosphate? **Ans.** Rb_3PO_4

 Logical