## Answers to:

a) Practice Problems 1 and 2 in page 239 and
b) Practice Problem 3 and 4 in page 240

Answers to Practice Problems F

1. a. $259.80 \mathrm{~g} / \mathrm{mol}$
b. $136.06 \mathrm{~g} / \mathrm{mol}$
c. $342.34 \mathrm{~g} / \mathrm{mol}$
d. $253.80 \mathrm{~g} / \mathrm{mol}$
e. $60.06 \mathrm{~g} / \mathrm{mol}$
f. $262.84 \mathrm{~g} / \mathrm{mol}$
2. a. $\mathrm{NaHCO}_{3}, 84.01 \mathrm{~g} / \mathrm{mol}$
b. $\mathrm{CeB}_{6}, 204.98 \mathrm{~g} / \mathrm{mol}$
c. $\mathrm{Mg}\left(\mathrm{ClO}_{4}\right)_{2}, 223.20 \mathrm{~g} / \mathrm{mol}$
d. $\mathrm{Al}_{2}\left(\mathrm{SO}_{4}\right)_{3}, 342.17 \mathrm{~g} / \mathrm{mol}$
e. $\mathrm{Fe}(\mathrm{OH})_{3}, 106.88 \mathrm{~g}$ gimol
f. $\mathrm{SnCl}_{2}, 189.61 \mathrm{~g} / \mathrm{mol}$
g. $\mathrm{P}_{4} \mathrm{O}_{10}, 283.88 \mathrm{~g} / \mathrm{mol}$
h. $\mathrm{ICl}, 162.35 \mathrm{~g} / \mathrm{mol}$
3. a. $92.15 \mathrm{~g} / \mathrm{mol}$
b. $0.0815 \mathrm{~mol} \mathrm{C}_{6} \mathrm{H}_{5} \mathrm{CH}_{3}$
4. a. $300.06 \mathrm{~g} / \mathrm{mol}$
b. $2.050 \mathrm{~g} \mathrm{PtCl} 2\left(\mathrm{NH}_{3}\right)_{2}$
